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every way. We cannot continue the observation of the solubility of the decahydrate beyond 32383, because it immediately splits up either into a less hydrated form--e.g. Na2S04.7H20--or the anhydrous form, Na2S04. The solubility curve of the heptahydrate meets the solubility curve of the anhydrous sulphate in the region of instability; transition point from the heptahydrate to the anhydrous salt is 34, or Na2S04.7H2ONa2S04+7H20 The so-called eutectic points E and Ez will be discussed later, but since the transformation of the anhydrous salt into the hydrate takes an appreciable time, it is possible to measure the approximate solubility of the anhydrous salt below 328. This is indicated by the dotted line in the diagram. In saturated solutions of hydrates, a definite hydrate is in dynamic equilibrium with the solution; if the hydrate changes as shown by E. Demarcay's study (1883) of the hydrates of thorium sulphate, the maximum amount of a salt which can enter into solution depends on its temperature and on its state of hydration; the solubilities of the different hydrates of a salt are different, and at the transition temperature, there is a break in the continuity of the solubility curve. H. W. B. Roozeboom's studies of the hydrates of a number of salts show that the solubility curves of the different hydrates of a salt indicate the limits of their stability. The solubilities of the two sodium sulphates--anhydrous and decahydrate--are quite different. If...
"Titles of chemical papers in British and foreign journals" included in Quarterly journal, v. 1-12.
Practical Chemical Thermodynamics for Geoscientists covers classical chemical thermodynamics and focuses on applications to practical problems in the geosciences, environmental sciences, and planetary sciences. This book will provide a strong theoretical foundation for students, while also proving beneficial for earth and planetary scientists seeking a review of thermodynamic principles and their application to a specific problem. Strong theoretical foundation and emphasis on applications Numerous worked examples in each chapter Brief historical summaries and biographies of key thermodynamicists—including their fundamental research and discoveries Extensive references to relevant literature
A unique text presenting practical information on the topic of nucleation and crystal growth processes from metastable solutions and melts Nucleation and Crystal Growth is a groundbreaking text thatoffers an overview and description of the processes and phenomena associated with metastability of solutions and melts. The author—a noted expert in the field—puts the emphasis on low-temperature solutions that are typically involved in crystallization in a wide range of industries. The text begins with a review of the basic knowledge of solutions and the fundamentals of crystallization processes. The author then explores topics related to the metastable state of solutions and melts from the standpoint of three-dimensional nucleation and crystal growth. Nucleation and Crystal Growth is the first text that contains a unified description and discussion of the many processes and phenomena occurring in the metastable zone of solutions and melts from the consideration of basic concepts of structure of crystallization. This important text: Outlines an interdisciplinary approach to the topic and offers an essential guide for crystal growth practitioners in materials science, physics, and chemical engineering Contains a comprehensive content that details the crystallization processes starting from the initial solutions and melts, all the way through nucleation, to the final crystal products Presents a unique focus and is the first

book on understanding, and exploiting, metastability of solutions and melts in crystallization processes Written for specialists and researchers in the fields of materials science, condensed matter physics, and chemical engineering. Nucleation and Crystal Growth is a practical resource filled with hands-on knowledge of nucleation and crystal growth processes from metastable solutions and melts. A new method of analysis for cuprous and hydrogen ion concentration gives rapid results, with no interference between ions. A solubility curve for cuprous chloride as a function of hydrochloric acid concentration was worked out. (Author).

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